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1-[2,6-Dichloro-4-(trifluoromethyl)-phenyl]-5-iodo-4-trifluoromethylsulfinyl-1*H*-pyrazole-3-carbonitrile

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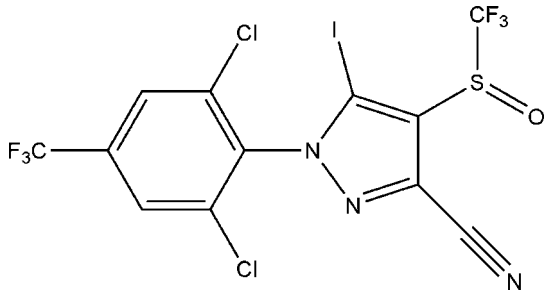
Key indicators: single-crystal X-ray study; T = 100 K; mean σ(C–C) = 0.007 Å; R factor = 0.041; wR factor = 0.111; data-to-parameter ratio = 13.3.

In the title compound, $C_{12}H_2Cl_2F_6IN_3OS$, the dihedral angle between the planes of the benzene and pyrazole rings is 77.8 (2)°. In the crystal, a short I···N contact of 2.897 (5) Å occurs.

Related literature

For related structures containing phenylpyrazole, see: Shi *et al.* (2009); Tang, Zhong, Li & Hu (2005). Tang, Zhong, Lin, Hu & Shi (2005).



Experimental

Crystal data

$C_{12}H_2Cl_2F_6IN_3OS$
M_r = 548.03
Monoclinic, $P2_1/c$
a = 9.5879 (3) Å
b = 13.7798 (4) Å
c = 14.3145 (5) Å
β = 107.400 (3)°
V = 1804.68 (10) Å^3
Z = 4
Cu $K\alpha$ radiation
μ = 18.41 mm^{-1}
T = 100 K
0.45 × 0.27 × 0.19 mm

Data collection

Oxford Diffraction Gemini (Cu) X-ray Ultra diffractometer
Absorption correction: analytical (*CrysAlis RED*; Oxford Diffraction 2006)
T_{min} = 0.016, T_{max} = 0.213
8203 measured reflections
3132 independent reflections
3065 reflections with $I > 2\sigma(I)$
R_{int} = 0.041

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.041
$wR(F^2)$ = 0.111
S = 1.13
3132 reflections
235 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.77 e Å^{-3}
$\Delta\rho_{min}$ = −1.25 e Å^{-3}

Data collection: *CrysAlis CCD* (Oxford Diffraction 2006); cell refinement: *CrysAlis RED* (Oxford Diffraction 2006); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEX* (McArdle, 1995); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT2972).

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supplementary materials

Acta Cryst. (2009). E**65**, o1774 [**doi:10.1107/S1600536809025082**]

1-[2,6-Dichloro-4-(trifluoromethyl)phenyl]-5-iodo-4-trifluoromethylsulfinyl-1*H*-pyrazole-3-carbonitrile

D.-X. Jiang and H.-H. Xu

Comment

The molecular structure of title compound, $C_{12}H_2Cl_2F_6IN_3OS$, is shown in Fig.1. The dihedral angle between the benzene ring and the pyrazole ring is 77.8 (2)°, while the corresponding ones in the two related compounds,$C_{12}H_4Cl_2F_6N_4S$ (Tang, Zhong, Li & Hu, 2005) and $C_{22}H_8Cl_4F_6N_8S_2$ (Tang, Zhong, Lin, Hu & Shi, 2005), are 83.2 (1)° and 88.2 (1)°, respectively.

Experimental

98% Fipronil (4.4 g, 10 mmol) was resolved in 40 ml chloroform in a 100 ml round bottom flask equipped with magnetic stirrer and a calcium chloride tube. Then iodine (3.6 g,14 mmol) was added into the solution. Half an hour later *tert*-butyl nitrite(1.43 g) was added into the solution and the mixture was heated under reflux for 2 h. Then it was left at room temperature overnight. The reaction mixture was filtered. Then the filtrate was evaporated *in vacuo*. The solid residue was purified by chromatography eluting using petroleum/ethyl acetate (4:1) and further recrystallized from toluene/hexane to afford colourless crystals. Yield: 4.70 g (86%).

Refinement

H atoms were included in calculated positions and treated in the subsequent refinement as riding atoms, with C—H = 0.93 Å and U_{iso}(H) = 1.2U_{eq}(C).

Figures

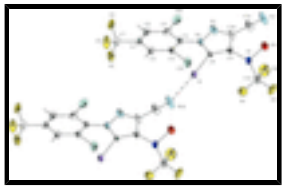


Fig. 1. The molecular structure of the title compound with the atom-numbering scheme. Displacement ellipsoids are drawn at the 30% probability level and H atoms are shown as small spheres of arbitary radii.

1-[2,6-Dichloro-4-(trifluoromethyl)phenyl]-5-iodo-4-trifluoromethylsulfinyl- 1*H*-pyrazole-3-carbonitrile

Crystal data

$C_{12}H_2Cl_2F_6IN_3OS$	F_{000} = 1040
M_r = 548.03	D_x = 2.017 Mg m^{-3}
Monoclinic, $P2_1/c$	Cu $K\alpha$ radiation, λ = 1.54184 Å
Hall symbol: -P 2ybc	Cell parameters from 3132 reflections
a = 9.5879 (3) Å	θ = 4.6–67.0º

b = 13.7798 (4) Å	μ = 18.41 mm^{-1}
c = 14.3145 (5) Å	T = 100 K
β = 107.400 (3)º	Prism, colorless
V = 1804.68 (10) Å^3	0.45 × 0.27 × 0.19 mm
Z = 4	

Data collection

Oxford Diffraction Gemini (Cu) X-ray Ultra diffractometer	3132 independent reflections
Radiation source: fine-focus sealed tube	3065 reflections with $I > 2\sigma(I)$
Monochromator: mirror	R_{int} = 0.041
T = 100 K	θ_{max} = 67.0º
ω and π scans	θ_{min} = 4.6º
Absorption correction: analytical (CrysAlis RED; Oxford Diffraction 2006)	$h = -10 \rightarrow 11$
T_{min} = 0.016, T_{max} = 0.213	$k = -16 \rightarrow 15$
8203 measured reflections	$l = -17 \rightarrow 16$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)]$ = 0.041	H-atom parameters constrained
$wR(F^2)$ = 0.111	$w = 1/[\sigma^2(F_o^2) + (0.0636P)^2 + 4.9895P]$ where $P = (F_o^2 + 2F_c^2)/3$
S = 1.13	$(\Delta/\sigma)_{max}$ = 0.001
3132 reflections	$\Delta\rho_{max}$ = 0.77 e Å^{-3}
235 parameters	$\Delta\rho_{min}$ = −1.25 e Å^{-3}
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}

I1	0.32209 (3)	0.65606 (2)	0.31318 (2)	0.02325 (15)
Cl1	0.49232 (14)	0.42087 (8)	0.14823 (9)	0.0336 (3)
Cl2	0.37074 (16)	0.80256 (8)	0.09630 (9)	0.0364 (3)
S1	0.70639 (13)	0.70720 (8)	0.47083 (8)	0.0257 (3)
O1	0.8422 (4)	0.7630 (3)	0.4932 (3)	0.0360 (8)
N1	0.5348 (4)	0.6307 (3)	0.1980 (3)	0.0215 (8)
N2	0.6758 (4)	0.6311 (4)	0.1956 (3)	0.0312 (10)
N3	1.0327 (6)	0.6582 (4)	0.3373 (4)	0.0474 (15)
C1	0.1794 (5)	0.5669 (3)	−0.0465 (3)	0.0238 (10)
C2	0.2659 (5)	0.4911 (3)	0.0021 (3)	0.0230 (9)
H2A	0.2448	0.4273	−0.0183	0.028*
C3	0.3855 (5)	0.5134 (3)	0.0823 (3)	0.0219 (9)
C4	0.4171 (5)	0.6096 (3)	0.1111 (3)	0.0217 (9)
C5	0.3304 (6)	0.6833 (4)	0.0603 (4)	0.0263 (10)
C6	0.2084 (6)	0.6628 (3)	−0.0189 (4)	0.0279 (11)
H6A	0.1480	0.7124	−0.0523	0.034*
C7	0.0517 (6)	0.5443 (4)	−0.1351 (4)	0.0304 (11)
C8	0.5210 (5)	0.6523 (3)	0.2870 (4)	0.0182 (9)
C9	0.6611 (5)	0.6671 (3)	0.3484 (4)	0.0228 (10)
C10	0.7517 (6)	0.6527 (3)	0.2873 (4)	0.0276 (11)
C11	0.9089 (6)	0.6571 (4)	0.3153 (5)	0.0350 (14)
C12	0.7590 (6)	0.5852 (4)	0.5261 (4)	0.0391 (13)
F1	0.0028 (4)	0.4545 (3)	−0.1333 (3)	0.0488 (9)
F2	0.0883 (4)	0.5527 (3)	−0.2176 (2)	0.0439 (8)
F3	−0.0603 (4)	0.6042 (3)	−0.1430 (3)	0.0487 (9)
F4	0.6443 (4)	0.5270 (2)	0.4989 (3)	0.0503 (9)
F5	0.7988 (5)	0.5949 (3)	0.6227 (3)	0.0585 (11)
F6	0.8683 (4)	0.5472 (3)	0.5007 (3)	0.0624 (11)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
I1	0.0165 (2)	0.0296 (2)	0.0220 (2)	−0.00092 (10)	0.00320 (13)	−0.00227 (10)
Cl1	0.0314 (6)	0.0221 (6)	0.0387 (7)	0.0056 (5)	−0.0026 (5)	0.0011 (5)
Cl2	0.0574 (8)	0.0171 (6)	0.0285 (6)	−0.0045 (5)	0.0033 (5)	−0.0014 (4)
S1	0.0264 (6)	0.0222 (6)	0.0230 (6)	−0.0021 (4)	−0.0008 (4)	−0.0002 (4)
O1	0.038 (2)	0.0290 (19)	0.036 (2)	−0.0106 (16)	0.0032 (16)	−0.0038 (16)
N1	0.0142 (18)	0.0247 (19)	0.0225 (19)	−0.0027 (16)	0.0008 (15)	−0.0014 (16)
N2	0.025 (2)	0.034 (2)	0.035 (2)	−0.0073 (17)	0.010 (2)	−0.0099 (19)
N3	0.021 (3)	0.064 (4)	0.058 (3)	−0.011 (2)	0.013 (2)	−0.035 (3)
C1	0.025 (2)	0.024 (2)	0.021 (2)	−0.0014 (19)	0.0050 (19)	−0.0004 (19)
C2	0.024 (2)	0.019 (2)	0.024 (2)	−0.0002 (18)	0.0043 (19)	−0.0038 (18)
C3	0.022 (2)	0.020 (2)	0.023 (2)	−0.0009 (18)	0.0050 (18)	−0.0004 (18)
C4	0.019 (2)	0.023 (2)	0.021 (2)	−0.0055 (18)	0.0039 (18)	−0.0053 (18)
C5	0.040 (3)	0.018 (2)	0.020 (2)	−0.002 (2)	0.007 (2)	−0.0022 (19)
C6	0.034 (3)	0.022 (2)	0.024 (2)	0.0029 (19)	0.002 (2)	0.0020 (18)
C7	0.025 (2)	0.031 (3)	0.030 (3)	0.003 (2)	0.000 (2)	−0.004 (2)
C8	0.010 (2)	0.018 (2)	0.025 (2)	0.0004 (15)	0.0038 (18)	0.0011 (16)

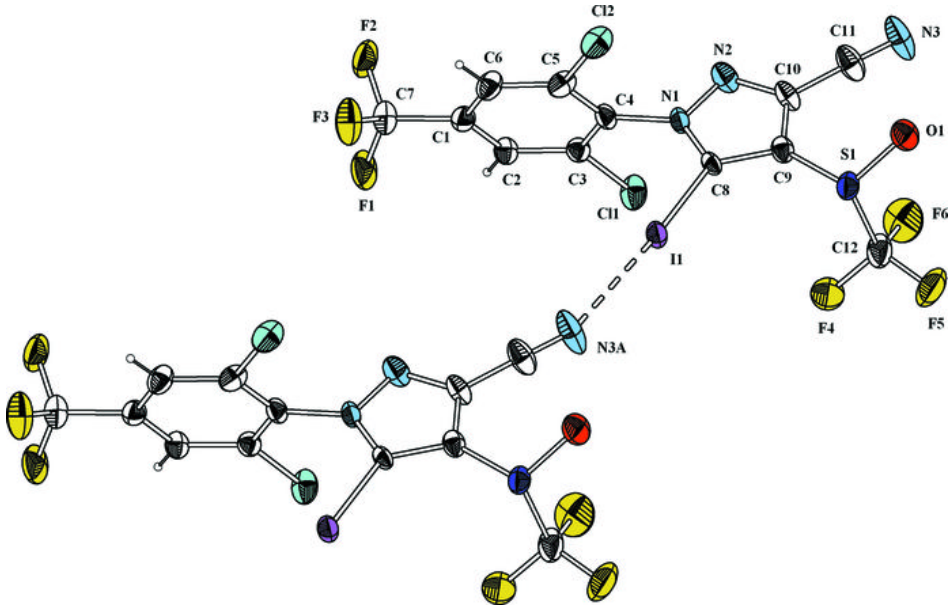
C9	0.021 (2)	0.017 (2)	0.027 (3)	−0.0013 (17)	0.002 (2)	−0.0014 (17)
C10	0.019 (3)	0.027 (3)	0.036 (3)	−0.0075 (18)	0.007 (2)	−0.009 (2)
C11	0.025 (3)	0.038 (3)	0.043 (3)	−0.008 (2)	0.012 (2)	−0.022 (2)
C12	0.034 (3)	0.029 (3)	0.042 (3)	−0.002 (2)	−0.008 (2)	0.009 (2)
F1	0.047 (2)	0.0386 (18)	0.0447 (19)	−0.0169 (16)	−0.0112 (15)	−0.0053 (15)
F2	0.0371 (18)	0.069 (2)	0.0202 (14)	0.0001 (16)	−0.0001 (12)	−0.0055 (15)
F3	0.0321 (17)	0.055 (2)	0.048 (2)	0.0130 (16)	−0.0052 (15)	−0.0163 (16)
F4	0.056 (2)	0.0313 (17)	0.053 (2)	−0.0147 (16)	0.0008 (17)	0.0124 (15)
F5	0.070 (3)	0.054 (2)	0.0337 (18)	−0.0074 (19)	−0.0127 (17)	0.0158 (16)
F6	0.047 (2)	0.049 (2)	0.084 (3)	0.0264 (18)	0.009 (2)	0.021 (2)

Geometric parameters (Å, °)

I1—C8	2.051 (4)	C2—H2A	0.9300
Cl1—C3	1.728 (5)	C3—C4	1.394 (6)
Cl2—C5	1.731 (5)	C4—C5	1.375 (7)
S1—O1	1.463 (4)	C5—C6	1.392 (7)
S1—C9	1.764 (5)	C6—H6A	0.9300
S1—C12	1.863 (6)	C7—F1	1.327 (6)
N1—C8	1.353 (6)	C7—F3	1.332 (6)
N1—N2	1.363 (6)	C7—F2	1.333 (6)
N1—C4	1.438 (6)	C8—C9	1.382 (7)
N2—C10	1.330 (7)	C9—C10	1.419 (7)
N3—C11	1.134 (8)	C10—C11	1.440 (8)
C1—C2	1.384 (7)	C12—F6	1.316 (8)
C1—C6	1.384 (7)	C12—F4	1.322 (7)
C1—C7	1.508 (6)	C12—F5	1.327 (7)
C2—C3	1.393 (7)		
O1—S1—C9	108.7 (2)	C5—C6—H6A	120.8
O1—S1—C12	105.8 (2)	F1—C7—F3	107.4 (4)
C9—S1—C12	95.4 (2)	F1—C7—F2	106.6 (4)
C8—N1—N2	113.6 (4)	F3—C7—F2	106.9 (4)
C8—N1—C4	126.0 (4)	F1—C7—C1	112.2 (4)
N2—N1—C4	120.5 (4)	F3—C7—C1	112.0 (4)
C10—N2—N1	103.3 (4)	F2—C7—C1	111.4 (4)
C2—C1—C6	122.5 (5)	N1—C8—C9	106.3 (4)
C2—C1—C7	118.7 (4)	N1—C8—I1	122.4 (3)
C6—C1—C7	118.7 (5)	C9—C8—I1	131.3 (4)
C1—C2—C3	118.0 (4)	C8—C9—C10	104.2 (4)
C1—C2—H2A	121.0	C8—C9—S1	125.6 (4)
C3—C2—H2A	121.0	C10—C9—S1	129.9 (4)
C2—C3—C4	120.4 (4)	N2—C10—C9	112.6 (5)
C2—C3—Cl1	119.7 (4)	N2—C10—C11	120.1 (5)
C4—C3—Cl1	119.9 (4)	C9—C10—C11	127.2 (5)
C5—C4—C3	120.2 (4)	N3—C11—C10	178.4 (6)
C5—C4—N1	120.1 (4)	F6—C12—F4	109.8 (5)
C3—C4—N1	119.6 (4)	F6—C12—F5	108.9 (5)
C4—C5—C6	120.6 (4)	F4—C12—F5	108.8 (5)
C4—C5—Cl2	119.9 (4)	F6—C12—S1	112.3 (4)

C6—C5—Cl2	119.5 (4)	F4—C12—S1	109.1 (4)
C1—C6—C5	118.3 (5)	F5—C12—S1	107.9 (4)
C1—C6—H6A	120.8		
C8—N1—N2—C10	0.7 (6)	C2—C1—C7—F2	−94.0 (5)
C4—N1—N2—C10	−179.8 (4)	C6—C1—C7—F2	83.8 (6)
C6—C1—C2—C3	0.7 (7)	N2—N1—C8—C9	−0.7 (5)
C7—C1—C2—C3	178.4 (4)	C4—N1—C8—C9	179.9 (4)
C1—C2—C3—C4	−0.9 (7)	N2—N1—C8—I1	−179.5 (3)
C1—C2—C3—Cl1	177.4 (4)	C4—N1—C8—I1	1.1 (6)
C2—C3—C4—C5	−0.2 (7)	N1—C8—C9—C10	0.3 (5)
Cl1—C3—C4—C5	−178.5 (4)	I1—C8—C9—C10	178.9 (3)
C2—C3—C4—N1	175.7 (4)	N1—C8—C9—S1	174.4 (3)
Cl1—C3—C4—N1	−2.6 (6)	I1—C8—C9—S1	−6.9 (6)
C8—N1—C4—C5	75.7 (6)	O1—S1—C9—C8	−149.2 (4)
N2—N1—C4—C5	−103.6 (5)	C12—S1—C9—C8	102.1 (4)
C8—N1—C4—C3	−100.2 (6)	O1—S1—C9—C10	23.4 (5)
N2—N1—C4—C3	80.4 (6)	C12—S1—C9—C10	−85.3 (5)
C3—C4—C5—C6	1.6 (7)	N1—N2—C10—C9	−0.5 (6)
N1—C4—C5—C6	−174.3 (5)	N1—N2—C10—C11	178.0 (5)
C3—C4—C5—Cl2	−179.8 (4)	C8—C9—C10—N2	0.2 (6)
N1—C4—C5—Cl2	4.3 (6)	S1—C9—C10—N2	−173.6 (4)
C2—C1—C6—C5	0.7 (8)	C8—C9—C10—C11	−178.3 (5)
C7—C1—C6—C5	−177.1 (5)	S1—C9—C10—C11	7.9 (8)
C4—C5—C6—C1	−1.8 (8)	O1—S1—C12—F6	−50.8 (5)
Cl2—C5—C6—C1	179.6 (4)	C9—S1—C12—F6	60.4 (4)
C2—C1—C7—F1	25.4 (6)	O1—S1—C12—F4	−172.7 (4)
C6—C1—C7—F1	−156.8 (5)	C9—S1—C12—F4	−61.5 (5)
C2—C1—C7—F3	146.3 (5)	O1—S1—C12—F5	69.2 (5)
C6—C1—C7—F3	−35.9 (7)	C9—S1—C12—F5	−179.6 (4)

Fig. 1


F2
Cl2
C11
N3
C6
C5
N2
C10
C7
C4
N1
F3
C1
O1
S1
C2
C3
C9
C8
F1
Cl1
F6
I1
C12
F4
F5
N3A